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Colligative properties

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are those that depend on the concentration of particles in a solution, not upon the identity of those particles.

Properties of Solutions - ScienceGeek.net

unsaturated solution. contains less solute than the solvent has the capacity to dissolve at a specific temperature. A . supersaturated solution. contains more

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solute than is present in a saturated solution at a specific temperature. Sodium acetate crystals rapidly form when a seed crystal is added to a supersaturated solution of sodium acetate.

PowerPoint Presentation - Chapter 13 Properties of Solutions

Determine its solubility

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when pressure of $O_2 = 1.00 \text{ atm}$. Units of Concentration Physical properties of solutions are often related to the concentration of the solute in the solution. Molarity Mole fraction: The same quantity we have used in fractional abundances as well as with gases (Dalton's law).

**PowerPoint
Presentation**

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Title: Colligative Properties of Solutions
1 Colligative Properties of Solutions. Boiling Pt Elevation and Freezing Pt Depression; 2 Colligative Properties. depend only on the number of solute particles in a solution ; does not depend on the identity of particles; 3 Boiling Point Elevation. a nonvolatile solute lowers the vapor pressure

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PPT - Colligative Properties of Solutions PowerPoint ...

Colligative properties of solutions -
Colligative properties of solutions The Effects of Solutes on Solvents | PowerPoint PPT presentation | free to view Chapter 19: Molality and Colligative Properties - Chapter 19: Molality and Colligative Properties

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HW Ch. 19 Blue Book:
#1-17, 19 (on
problems that are a-z,
please do a and b only)
Use the ...

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Properties of Solutions - SlideShare

COLLIGATIVE

PROPERTIES Elevation of Boiling Point

Depression of Freezing Point Lowering of

Vapor Pressure

Osmotic Pressure

MOLE FRACTION &

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MOLALITY MOLE
FRACTION OF

Component $i = X_i = n_i$
/ n_{total} (c.f Gases;
Chapter 5, p.217)

MOLALITY = Moles of
Solute / kg Solvent

MOLALITY Useful when
Temperature Changes
are considered, as
Volumes of solutions
change with changing
temperature, whereas
...

**COLLIGATIVE
PROPERTIES**

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Chapter 11 - Properties of Solutions . 11.1

Solution Composition .

A. Molarity 1. liters of solution moles solute

Molarity(M) = B. Mass

Percent 1. $\times 100 =$

mass of. solution mass of solute Mass percent.

C. Mole Fraction . 1. D.

Molality 1. ki ram of solvent moles of solute

Molality log = E.

Normality 1. liter of solution equivalents

Chapter 11 -
Page 13/27

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Properties of Solutions

Concentration of Solution Solvent Solute

Molarity Parts ratio

Mole Fraction Molality

Concentration of

Solution Moles of

solute Liter of solution

(M) = = Mol L amount

of solute (g or ml)

amount of solution (g

or ml) (102) or (106) or

(109) Moles of solute

Total moles of solution

(c) = = Kilograms of

solvent Moles of solute

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(m) = Molarity Molarity
Example Problem 1
12.6 g of NaCl are
dissolved ...

PowerPoint Presentation

1. Solutions Colligative Properties • Changes in colligative properties depend only on the number of solute particles present, not on the identity of the solute particles. • Among colligative properties are □ Vapor

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pressure lowering
□ Boiling point elevation
□ Melting point depression □ Osmotic Pressure.

Colligative properties - SlideShare

Colligative properties of a solution depend on only the total number of dissolved particles in solution, not on their chemical identity.

Colligative properties include vapor pressure,

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boiling point, freezing point, and osmotic pressure. The addition of a nonvolatile solute (one without a measurable vapor pressure)...

13: Properties of Solutions - Chemistry LibreTexts

Basic Chemistry
Tutorial: Properties of Solutions . Shane Plunkett .
plunkes@tcd.ie - Solids

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- Structure of solids - Liquids • Vapour pressure - Solutions • Solubility of gases in liquids • Henry's law, Le Chatelier's principle • Solubility of liquids in liquids • Vapour pressure of solutions • Colligative properties

Basic Chemistry Tutorial: Properties of Solutions

five senses description
Examples of Physical
Properties Color Smell

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Taste Hardness State
of Matter Boiling,
Freezing, or Melting
Point Examples of
Physical Properties
Density Mass Volume
Malleability (the ability
to be molded)
Solubility (the ability to
be dissolved) Chemical
Properties Are
determined by a
substance's ability to
with other ...

**Physical and
Chemical Changes**

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and Properties

EQ: What are Mixtures and Solutions? This is a bundled lesson that includes the student interactive notes and the Power Point presentation. Student discussions (Turn and Talks) embedded in presentation. Students will differentiate between mixtures and solutions. Students will identify or describe the factors that affect solubility.

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What are Mixtures and Solutions? PowerPoint and Student ...

Electrical Properties of Colloidal Solutions The particles of the colloidal solution carry the same type of charge, while the dispersion medium carries an equal and opposite charge. The charge on the dispersion medium balances the charge on

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dispersed particles and the solution as a whole is electrically neutral.

Properties of Colloidal Solutions: Physical, Optical ...

Colligative Properties-
Page 1. Lecture 4:

Colligative Properties •

By definition a colligative property is a solution property (a property of mixtures) for which it is the amount of solute dissolved in the solvent

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matters but the kind of solute does not matter.

Colligative Properties- Page 1 Lecture 4: Colligative ...

ppt on colligative properties of solutions are a good way to achieve details about operating certain products. Many products that you buy can be obtained using instruction manuals.

These user guides are

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clearly built to give step-by-step information about how you ought to go ahead in

PPT ON COLLIGATIVE PROPERTIES OF SOLUTIONS PDF

Slide 1. Colligative Properties of Solutions. Jacobus Henricus van 't Hoff (1852-1911) Slide

2. Colligative Properties. Colligative properties are those that depend on the concentration of

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particles in a solution, not upon the identity of those particles.

Colligative Properties of Solutions - Presentation Chemistry

Concentration of Solutions and Molarity
The concentration of a solution is a measure of the amount of solute that is dissolved in a given quantity of solvent. -A dilute

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solution is one that contains a small amount of solute. -A concentrated solution contains a large amount of solute.

Concentration of Solutions and Molarity

alpha.chem.umb.edu

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